TABLE I^{a}

Fig.	Compound (Solvent)	Long-wave length region			Short-wave length region		
		$[\alpha]^{b} \times 10^{-4}$	$[\theta]_{max} \times 10^{-4}$	$\epsilon_{max} \times 10^{-2}$	$[\alpha]^{b} \times 10^{-4}$	$[\theta]_{max} \times 10^{-4}$	$\epsilon_{maq} \times 10^{-2}$
1	I (isoöctane)	+14.2 (304)^{c}	+ 9.85 (307)	8.05 (308)^{j}	...	Positive to 250 mμ	66 (228)
2	II (isoöctane)	+ 7.8 (302)^{d,e}	+ 3.18 (306)^{e}	2.77 (307.5)^{c}	...	Positive below 250 mμ	...
3	III (isoöctane)	+ 5.2 (298)^{d,f}	+ 3.35 (296)^{f}	1.22 (296)^{c}	...	...	...
4	IV (isoöctane)	+13.7 (297)^{g}	+10.8 (296)	6.70 (297.5)^{i}	...	+9.65 (245)	98 (245)^{g}
4	V (isoöctane)	−15.2 (301)^{g,h}	−10.6 (297)	6.70 (297.5)^{i}	−16.2 (243)^{g,h}	−9.5 (245)	98 (245)^{g}
5	VI (isoöctane)	...	− 0.33 (282) + 0.60 (269)	6.0 (274)^{g,i}	−13.0 (242)^{g}	−7.7 (240)	115 (240.5)^{g}
5	VII (dioxane)	...	+ 0.59 (271)	8.90 (281)^{g,i} 19.3 (270)^{g}	−9.8 (241)^{g}	−7.8 (245)	111 (244.5)^{g}
6	VIII (dioxane)	...	− 4.15 (330)	142 (348)^{g}	+157 (253)^{g}	+8.5 (255)	519 (240)^{g}

^{a} Values in parentheses are wave lengths in mμ. ^{b} Molecular amplitude of O.R.D. curves with wave lengths taken halfway between extrema. ^{c} A. Moscowitz, K. Mislow, M. A. W. Glass and C. Djerassi, *J. Am. Chem. Soc.*, **84**, 1945 (1962). ^{d} K. Mislow and J. G. Berger, *ibid.*, **84**, 1856 (1962). ^{e} After correction to optical purity assuming 48% optical purity for the sample (ref. *d*). ^{f} After correction to optical purity assuming 40% optical purity for the sample (ref. *d*). ^{g} K. Mislow, M. A. W. Glass, R. E. O'Brien, P. Rutkin, D. H. Steinberg, J. Weiss and C. Djerassi, *J. Am. Chem. Soc.*, **84**, 1455 (1962). ^{h} Refers to solvent dioxane. ^{i} Shoulder. ^{j} After subtraction of the aromatic absorption band (ref. *c*).

Acetals and Ketals with Lithium Aluminum Hydride—Aluminum Chloride.

Page 2372. In col. 2, Table III, entry 5, the bond should be to the methine carbon rather than the methyl carbon at the bottom of the formula.

Page 2374. In col. 2, last line, for "lit.^{63}" read "lit.^{33}."

Page 2375. In col. 1, line 15, for "α-hydroxypropyl" read "γ-hydroxypropyl."

Page 2376. In col. 1, second heading, for "2-Cyclohexyloxy-2-methyl-2-propanol" read "2-Cyclohexyloxy-2-methyl-1-propanol." In col. 2, line 26, end of line for "n^{20}D" read "n^{25}D."

Page 2377. In Table IV, the entries given for Vb, VIb and VIIb under *m/e* 27 actually refer to *m/e* 28. The correct entries for *m/e* 27 are 36.06, 36.03 and 35.96. In col. 2, line 5, for "94° (12 mm.)" read "84° (12 mm.)."—E. L. ELIEL.

Ernest L. Eliel, Louis A. Pilato and Victor G. Badding. Reduction with Metal Hydrides. XIII. Hydrogenolysis of Hemithioacetals and Hemithioketals with Lithium Aluminum Hydride—Aluminum Chloride.

Page 2381. In col. 2, line 17, for "284.0 g." read "284.0 mg." Line 35–36, for "m.p. 118–114°" read "m.p. 118–144°."

Page 2382. In Table III, col. 1, first heading, for "Ethylene ketal- or hemithio acetal of" read "Ethylene hemithioketal or hemithioacetal of." Second heading, for "Trimethylene acetal or hemithioketal of" read "Trimethylene hemithioacetal or hemithioketal of."—E. L. ELIEL.

John L. Kice and Kerry W. Bowers. Mechanism of Reactions of Sulfinic Acids. II.

Page 2386. In eq. 5b for "k_o" read "k_0." In eq. 5c for "k_0" read "k_e."—JOHN L. KICE.

R. L. Hinman and E. B. Whipple. The Protonation of Indoles: Position of Protonation.

Page 2535. To correct for errors in the original calculations of log ε the ultraviolet absorption curves of Fig. 1 should be translated vertically by these amounts: curve 1, no change; curve 2, −0.15; curve 3, −0.38; curve 4, −0.15; curve 5, no change.

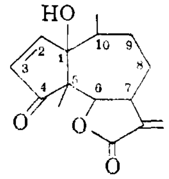
Page 2538. The data from the ultraviolet absorption curves of 2,3,3-trimethylindolenine should read:

	λ_{max}, mμ			ϵ_{max}		
In 95% ethanol	222.5	255		11500	6220	
In 0.1 *N* HCl	229	235	275	7100	6150	5840

RICHARD L. HINMAN

W. Herz, H. Watanabe, M. Miyazaki and Y. Kishida. The Structures of Parthenin and Ambrosin.

Page 2603. Formula I should be



WERNER HERZ.

E. Bunnenberg, Carl Djerassi, Kurt Mislow and Albert Moscowitz. Correspondence between Absorption and Circular Dichroism Curves of Inherently Dissymmetric Chromophores.

Pages 2823 ff. In our first communication on the subject we alluded to the close relationship which exists between absorption, optical rotatory dispersion and circular dichroism (C.D.) curves of optically active substances possessing inherently dissymmetric chromophores. At that time we remarked on the observation that the long-wave length C.D. bands were somewhat narrower than the corresponding absorption bands and that the short-wave length C.D. maxima were red-shifted relative to the absorption bands. Although these displacements did not seem to be unreasonable, we have now found that *the C.D. and absorption curves are actually in much closer correspondence* (see corrected portions of Table I above to be substituted in original Table I). The earlier apparent discrepancies were largely artifacts resulting from mechanical failure in the wave length scan mechanism of the monochromator in the Baird-Atomic/Jouan Dichrograph (Model JO-1) which was used in the work.—CARL DJERASSI.

R. G. Hayter. Transition Metal Complexes of Secondary Phosphines. I. Palladium(II) and Diphenylphosphine.

Page 3049. In col. 2, line 2 from the end, for "1" read "0.586."

Page 3050. In Table IV, col. 3, for "1" in lines 1-3, read "0.586." In col. 5, for "0.862" in lines 1-3, read "1.47."—R. G. HAYTER.

Jay K. Kochi. Oxidation of Allylic Radicals by Electron Transfer: Effect of Complex Copper Salts.

Page 3271. In Fig. 1, the ordinate division numbers should read 50 to 100% instead of 40 to 90% as printed.—JAY K. KOCHI.

L. S. Forster, A. Moscowitz, J. G. Berger and K. Mislow. Inherently Dissymmetric Chromophores. Optical Activity Associated with a Homoconjugated Diene.

Page 4353 ff. Compound I designates (+)-(1R)-5-methylenebicyclo[2.2.1]hept-2-ene and not (+)-(1R)-5-methylenebicyclo[2.2.2]hept-2-ene as incorrectly indicated in line 3 of the text and in the caption of the figure.—ALBERT MOSCOWITZ.